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1. The protons in the nucleus do not change during normal chemical

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reactions. Only the outer electrons move.

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CHAPTER 7 REVIEW

Chemical Formulas and
Chemical Compounds

SECTION 2 SHORT

ANSWER Answer the
following questions in
the space provided. 1.

Assign the oxidation
number to the

specified element in
each of the following

examples: 4 a. S in H

2SO 3 6 b. S in MgSO 4

2 c. S in K 2S 1 d. Cu in

Cu 2S 6 e. Cr in Na

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Answers
2CrO₄ 5 f. N in HNO₃
4 g. C in (HCO₃)₃ 3 h. N
in (NH₄)₂ SCl₂ a.

7 Chemical Formulas and Chemical Compounds

Chemistry Chapter 7
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talking about

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Chemistry Chapter 7

Short Answer Type

Question 1, White

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phosphorus has structure as given below: It consists discrete tetrahedral P₄ molecules. 1. Red phosphorus has... 2. It is poisonous, insoluble in water but soluble in CS₂. 2. It is odourless, non - poisonous insoluble in water ...

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Ans: Chapter 7 of CBSE
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discusses the properties of the p-Block elements in a detailed manner. The elements from Group 13 to 18 in the periodic table are characterized as the p-Block elements. The study of the elements and their electronic configuration is an important part of Physical Chemistry.

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Chapter 7 The p ...**

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Class 11 Chemistry
Chapter 7 Very Short
Answer Type

Questions. Question 1.

What is meant by
equilibrium? Answer:
Equilibrium is a state at
which rate of
forwarding reaction is
equal to the rate of
backward reaction.

Question 2. State the
law of mass action?

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Chapter 7

Equilibrium

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Balancing Chemical

Equations Answer Key

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Balance the equations

below: 1) $1 \text{ N}_2 + 3 \text{ H}_2$

$\rightarrow 2 \text{ NH}_3$ 2) 2 KClO_3

$\rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) 2

$\text{NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF}$

$+ 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2$

$\rightarrow 2 \text{ H}_2\text{O}$ 5) 1 Pb(OH)

$2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} +$

1 PbCl_2 6) $2 \text{ AlBr}_3 + 3$

$\text{K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}$

$2(\text{SO}_4)_3$ 7) $1 \text{ CH}_4 + 2$

$\text{O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$

8) $1 \text{ C}_3\text{H}_8 + 5 \text{ O}_2 \rightarrow$

$3 \text{ CO}_2 + 4 \text{ H}_2\text{O}$ 9) 2 C

$8\text{H}_{18} + 25 \text{ O}_2 \rightarrow 16$

$\text{CO}_2 + 18 \text{ H}_2\text{O}$...

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Worksheet #1

Balancing Chemical Equations

Mastering Chemistry

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Answer Key Chapter 7 -

Chemistry: Atoms First

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equation is balanced

when the same

number of each

element is represented

on the reactant and

product sides.

Equations must be

balanced to accurately

reflect the law of

conservation of matter.

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Answer Key Chapter 7 - Chemistry: Atoms First | OpenStax

Answer Key Chapter 7 -
Chemistry: Atoms First
2e | OpenStax 1. An
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Chapter 7 Chem:
Practice Test

Questions and Answers

. 1. A chemical bond is the attraction between a positive nucleus and negative electrons, or positive and negative ions. ... 7. The number of valence electrons can be directly related to the difference in

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reactivity. A. True. B.

False. 8.

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