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CorrectionKey=NL-A

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be made ...**

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Stoichiometry 9.1

Objectives ... •9.3

Worksheet to STOP 1

•There is more practice
on the LR Problem Sets

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Yield •The theoretical

yield is the maximum

amount of product that

can be produced from

a given amount of

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reactant.

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Standard Review

Worksheet 1. Answers
will vary. An example is
included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$
This describes the
decomposition reaction
of hydrogen peroxide.

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CHAPTER 9 REVIEW

Stoichiometry. 1.

Stoichiometry Worksheet Answers Chapter 9

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical

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reaction. Reaction

stoichiometry is the subject of this chapter and it is based on

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CHAPTER 9 REVIEW

Stoichiometry SECTION

3 PROBLEMS Write the

answer on the line to

the left. Show all your

work in the space

provided. 1. 88% The

actual yield of a

reaction is 22 g and

the theoretical yield is

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25 g. Calculate the percentage yield. 2.

6.0 mol of N₂ are mixed with 12.0 mol of H₂ according to the following equation: N₂(g) + 3H₂(g) → 2NH₃(g)

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5.0 g Cu 1 mol Cu 1

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mol Ag 107.9 g Ag =

8.5 g Ag. 63.5 g Cu 1

mol Cu 1 mol Cu. $8.5 \times 100 = 55.9\%$ yield.

15.2 CHAPTER 11:
STOICHIOMETRY. MOLE
TO MOLE RATIO. When
nitrogen and hydrogen
gas are heated under
the correct conditions,
ammonia gas (NH₃) is
formed. a. RXN: 1. N₂
+ 3. H₂ (2. NH₃. b.

CHAPTER 11: STOICHIOMETRY

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Remember to convert
all temperatures to ...

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